Solutions Review

Homogenous mixture: also called aqueous, also called a solution. Evenly distributed.

NaCl (aq): NaCl is dissolved in water

Heterogeneous mixture: unevenly distributed. Ex. Sand in water

Separation of mixtures

Distillation: separation of 2 liquids by boiling point Chromatography: separation by density and polarity



 Filtration: separating a solid from a liquid (heterogeneous mixture)



 Evaporation: separating a dissolved solid from the liquid (homogenous mixture)

Polar vs. Non polar molecules:



Like dissolves like. Water is polar and will dissolve other polar substances, Non polar substances do not dissolve in water (ex oil)

Double replacement reactions:

Predicting products



 3 LiOH + Fe(NO3)3 → Fe(OH) 3+ 3 Li(NO3)

Fe(OH)3 would form a precipitate based on table F

 AgF + Na2CO3 →

Precipitate: \_\_\_\_\_\_\_\_\_\_\_\_\_\_

 NaOH + H3PO4 →

Precipitate: \_\_\_\_\_\_\_\_\_\_\_\_\_\_

Ba(CN)2 + (NH4)2SO4 →

Precipitate: \_\_\_\_\_\_\_\_\_\_\_\_\_\_